

STABILITY OF CLOSE-CELL Al FOAMS DEPENDING ON THE USAGE OF DIFFERENT FOAMING AGENTS

STABILNOST ALUMINIJEVIH PEN Z ZAPRTO POROZNOSTJO GLEDE NA UPORABO RAZLIČNIH PENILNIH SREDSTEV

Irena Paulin

Institute of Metals and Technology, Lepi pot 11, 1000 Ljubljana, Slovenia
irena.paulin@imt.si

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Close-cell Al foams produced by the powder-metallurgy (PM) route can be made with different foaming agents, regarding the type of liberated gas. Most commonly, H₂ gas is used as the liberation agent, but there is a huge improvement made with CO₂ gas liberating agents, such as calcite and dolomite. In order to determine the benefits and/or disadvantages of foaming agents, studies of the foam pores' stability, depending on the type of liberated gas, were performed. The stability of Al foams was studied by different analytical techniques, i.e., AES, expandometer, heating microscopy and SEM/EDS. AlSi12 aluminium powder as the matrix material and TiH₂ and CaCO₃ as the foaming agents liberating different gases – one based on H₂ and another on CO₂ – were used. Based on the obtained results, the mechanism of foam stability was studied and a comparison of the two foaming agents was made and evaluated.

Keywords: Al foams, CaCO₃, TiH₂, stability of pores, oxide layers

Aluminijeve pene z zaprto poroznostjo, narejene po postopku metalurgije prahov, se lahko pripravi z različnimi penilnimi sredstvi glede na plin, ki se pri penjenju sprošča. Najbolj pogosto je uporabljeno penilno sredstvo na osnovi H₂ plina, vendar se v zadnjem času veliko uporabljajo tudi penila na osnovi CO₂, kot sta kalcit in dolomit. Da bi ugotovili prednosti in slabosti različnih penilnih sredstev, so bile narejene študije stabilnosti por glede na tip izhajajočega plina. Stabilnost aluminijevih pen je bila raziskovana z različnimi analitskimi tehnikami – z AES, ekspandometrom, segrevalnim mikroskopom in SEM/EDS analizo. Za raziskavo smo uporabili aluminijev prah AlSi12 in TiH₂ ter CaCO₃ kot penilni sredstvi z različnimi tipoma sproščenega plina – en na osnovi H₂ in drugi na osnovi CO₂. Glede na rezultate raziskave smo opisali mehanizem za stabilizacijo aluminijevih pen in primerjali ter ocenili obe penilni sredstvi.

Ključne besede: Al pene, CaCO₃, TiH₂, stabilnost por, oksidne plasti

1 INTRODUCTION

Al foams as a promising class of materials with great chances for applicative use¹⁻³ due to their mechanical, chemical and physical properties were analyzed from different viewpoints. Most studies were performed on the synthesis, characterization and mechanical testing of foams, but few investigated the stability of foams by adding ceramic particles into the matrix and no analysis was made to explain the stability of the interior of the pores and the properties of the material depending on this stability.

The PM production process starts with mixing metal powders – elementary-metal powders, or alloyed powders or metal powder blends – and a foaming agent, after which the mixture is compacted to yield a dense, semi-finished product.^{4,5} The method depends on the preparation of the precursors⁶, which in principle can be made by any technique that ensures the embedding of the foaming agent into the metal matrix without any observable residual open porosity. In general, precursors consist of a compacted metallic powder and a foaming agent that are sintered at a pre-determined temperature. Due to the high temperature of the thermal treatment, the

foaming agent decomposes into a solid component that is incorporated into the matrix material, and a gas component that causes foaming of the matrix material.⁷

The metal matrix is a semi-solid, thus the gas liberated from the agent is forming pores that give a specific shape and mechanical properties to the foams. The obtained foam must be rapidly and properly cooled down to retain the bloated structure. Pore formation depends on the amount and the type of the liberated gas, the time and temperature of the thermal treatment and the cooling process. All those parameters influence the stability of the pores and thus the stability of the material properties. The quantitative expansion and collapsing behaviors of the samples were characterized with a mechanical expandometer and heating microscopy.

This investigation is focused on the closed-cell Al foam produced by the powder metallurgy (PM) process using different gas-liberating foaming agents, such as TiH₂ and CaCO₃. The aim of the study was to investigate the mechanism of the stability of the Al foams to understand the process of pore formation, the possibility of predicting the foam properties and for the optimal selection of the foaming agent, depending on the expected properties.

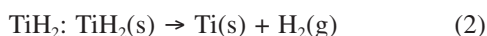
2 EXPERIMENTAL PROCEDURES

In the present study of Al-foam stability, the interior surfaces of the pores were carefully studied. AlSi12 aluminium powder, TiH₂ and CaCO₃ as the foaming agents liberating different gases – one based on H₂ and the other on CO₂ – were used. The decomposition of both foaming agents could be described by the two chemical reactions given below, and the volume of liberated gas was calculated from the ideal gas law:

$$pV = nRT \quad (1)$$

where p is the pressure, V is the volume of gas, n is the amount of substance ($n = m/M$, where m is the mass and M is the molar mass), R is the gas constant and T is the temperature.

The decomposition of:



1 g of TiH₂ liberates at $T = 923 \text{ K}$ (650 °C) a gas volume of $V = 1.51 \text{ L}$



1 g of CaCO₃ liberates at $T = 923 \text{ K}$ (650 °C) a gas volume of $V = 0.76 \text{ L}$

The calculation of the gas liberation at standard temperature and pressure shows that the volume of liberated gas is relatively high, but in the real foaming process, where the pressure of gas in the material is higher, the volume of gas is smaller. On the other hand, some of the liberated gas escapes out of the material when the pressure in the formed bubbles is too high.^{8,9} The process of pore formation can be explained by the theory of soap bubbles.

A calculation of the amount of liberated gas shows that the largest amount of gas liberated under the same conditions is obtained with the TiH₂ decomposition, while the amount of liberated gas with the decomposition of CaCO₃ is just a half of that. A calculation of the volumes of liberated gas gives how much foaming agent is needed for similar results of foaming.

All the samples were prepared by the same procedure, i.e., mixing the AlSi12 powder with mass fractions 1 % TiH₂ and 3 % CaCO₃, respectively, cold compaction at 1200 MPa⁶ into semi-products, called precursors, and foaming at a temperature that is specific for each foaming agent.¹⁰ After preparing the foams, some in a lab furnace and some in the expandometer, specimens were analyzed with different techniques: SEM/EDS (SEM JEOL 6500F with Oxford INCA EDX analyzer) was used for micro-chemical analyses of the inclusions in the foam thin walls, while the surface of the interior of the pores and the exterior of the foams were investigated by AES (Microlab 310 F VG–Scientific). AES depth profiles were performed by ion etching with a velocity of 0.125 nm/min. The AES surface analyses were also performed for the initial powder particles to confirm the presence of an oxide layer on the surface.

The mechanism of foam stability was explained by the investigation with heating microscopy and the expandometer. Both methods were used to determine the time and the temperature of foaming the material. In heating microscopy, small samples of precursor material with different foaming agents were put into the furnace within the microscope and heated up to 800 °C with a heating rate of 7 °C/min. Precursors were inserted at room temperature into the microscope and heated up until the foaming process did not stop and foams started to collapse. During the heating, changes to the shape and the size of the investigated material were observed. This method enabled us to determine the temperature of the foaming process start and the temperature of the foam collapse.

The mechanical expandometer was used to measure the expansion of the foamable precursors inside a cylindrical mould as a function of time and temperature.¹¹ The temperature of the expandometer furnace was held constant at 750 °C for the TiH₂ foaming agent and at 800 °C for the CaCO₃ foaming agent.

3 RESULTS AND DISCUSSION

The foaming process was studied by observing precursors in the heating microscope during heating. Samples began to change their shape and volume at a certain temperature, depending on the type of foaming agent. The results obtained with the heating microscope enabled us to determine the approximate time and temperature of the beginning of the process. Changes of the foam expansion (precursor was made with AlSi12 + TiH₂) are represented in **Figure 1**.

The heating microscope also enabled us to observe the collapsing of the foams after extended foaming times at individual temperatures. This resulted in gas escaping from the bubbles that were formed inside the matrix material. At a certain temperature, the matrix material became viscous, which prevented the escape of gas from the material. Higher temperatures caused a drop in the material viscosity, and due to pressure in the gas bubbles the gas escaped and the foam collapsed. That was the reason that the foams had to be cooled down fast after the desired expansion of material was achieved. Cooling could be achieved either by immersing the foamed product into water or with cold compressed air. In our case cooling in water was applied.

Foams prepared with different foaming agents were cut into smaller pieces and samples for the investigation with SEM/EDS were prepared, using a standard metallographic procedure with grinding and polishing. **Figures 2 and 3** show SE (secondary electron) and BE (back-scattered) images of the polished foamed material, where residuals of solid particles of foaming agents captured in the walls of the foams after the gas liberation are visible.

In^{12,13} the authors confirmed that those ceramic particles helped to stabilize the Al foams. In analyses of

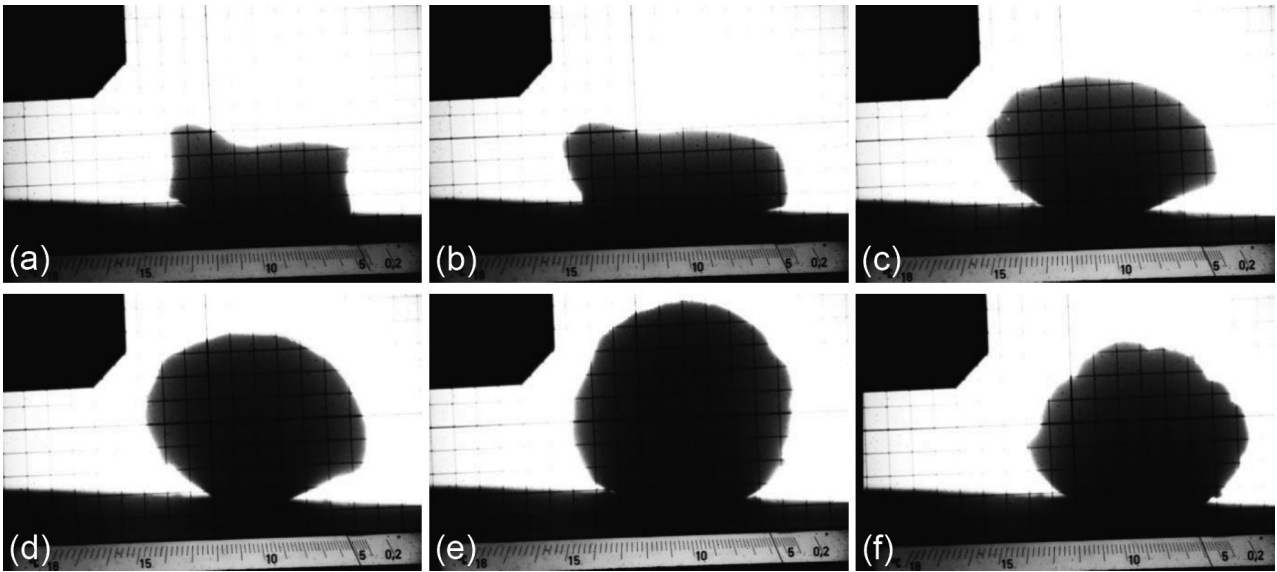


Figure 1: Heating microscope – change of precursor AlSi12 + TiH₂ size: a) 420 °C, no changes, b) 500 °C, first noticed change, c) 520 °C, precursor starts rapidly to bloat, d) 525 °C, still rapidly bloating, e) 580 °C, maximum expansion, f) 600 °C, collapse of sample, keeping the shape of a loaf

Slika 1: Segrevalna mikroskopija – sprememba velikosti prekurzorja iz AlSi12 + TiH₂: a) 420 °C, še nobenih sprememb velikosti, b) 500 °C, prve spremembe, c) 520 °C, vzorec začne hitro naraščati, d) 525 °C, vzorec zelo hitro narašča, e) 580 °C, vzorec doseže največjo velikost, f) 600 °C, vzorec se sesede, vendar obdrži obliko hlebčka

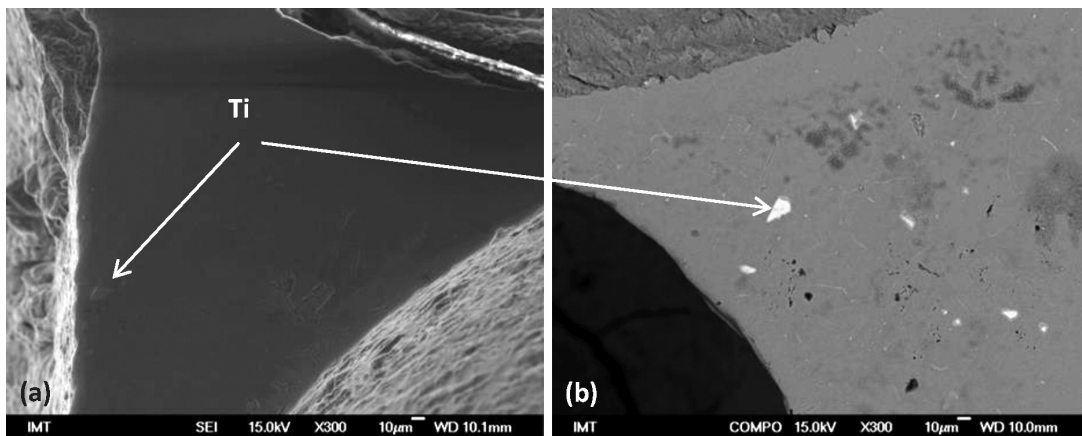


Figure 2: SEM images of thin foam-cell wall with Ti residuals of TiH₂ foaming agent: a) SE image and b) BE image

Slika 2: SEM posnetka tanke celične stene aluminijeve pene z ostanki Ti iz penilnega sredstva TiH₂: a) SE posnetek in b) BE posnetek

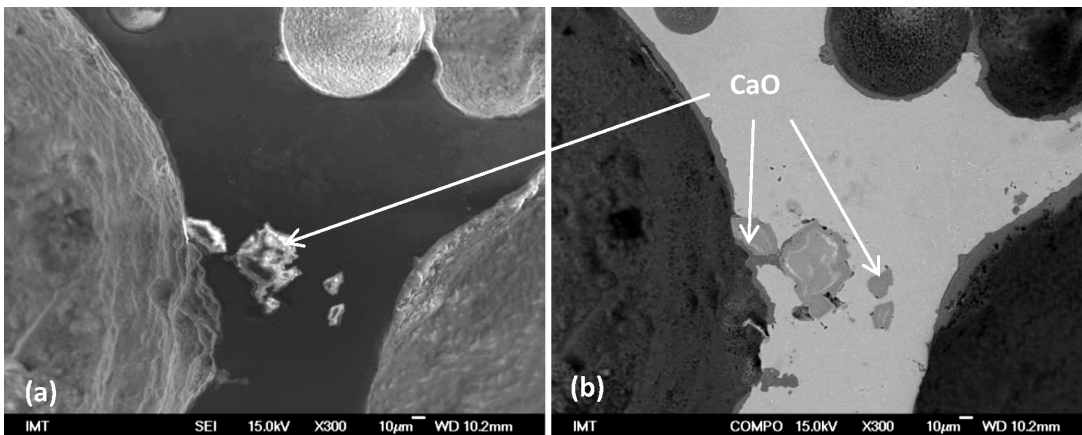


Figure 3: SEM images of thin foam-cell wall with residuals of CaCO₃ foaming agent: a) SE image and b) BE image; the same spot

Slika 3: SEM posnetka tanke celične stene z ostanki penila CaCO₃: a) SE posnetek in b) BE posnetek; isto mesto

ceramic additions to Al foams during preparation, a special emphasis was given to the positive effects of CaO particles.

One of the important properties of aluminium foams is also their stability during the foaming process. The quantitative expansion and collapse behavior of the samples were characterized with the mechanical expandometer, in which the precursor was exposed to a constant temperature of 750 °C for AlSi12 + TiH₂ and 800 °C for AlSi12 + CaCO₃. The expansion was measured with seven samples prepared using TiH₂ and CaCO₃ agents, respectively, and it was approximately 245 % with the CaCO₃ and up to 285 % with the TiH₂ precursors. Nevertheless, the TiH₂ precursors appeared to collapse faster than the CaCO₃ precursors.

Observations in the mechanical expandometer and in the heating microscope revealed that the foams prepared with a foaming agent based on the liberation of CO₂ gas were more stable than those that were prepared with the liberation of H₂. The obtained results were confirmed by the AES analysis of the interior surfaces of the pores where the oxide layers were analyzed and thus the stability can be explained.^{6,14,15} Two interior pore surfaces, obtained by foaming AlSi12 + TiH₂ and AlSi12 + CaCO₃ mixtures, were analyzed, and the results were compared (Figures 4 and 5). Also, the external surface of the foam was analyzed to see the difference in the interior and exterior surfaces of the foam (Figure 6).

The results showed differences in the thicknesses of the oxide layers on the surface of the pore interior depending on the applied foaming agent and on the surface of the foam exterior. The thickness of the oxide layer on the interior pore surface obtained by foaming with TiH₂ was 10–15 nm, while the thickness of the oxide layer on the external surface of the foam was 35–40 nm. On the other hand, the thickness of the oxide layer on the interior pore surface obtained by foaming with CaCO₃ was 90–125 nm, while the oxide layer on the external surface of the foam was about 45–60 nm thick. The difference in

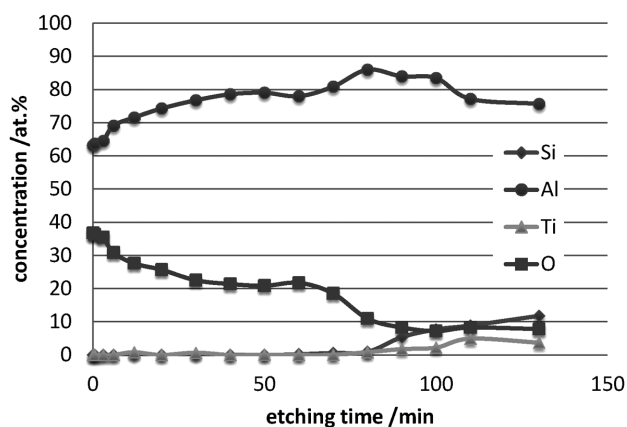


Figure 4: AES depth profiles of the interior surface of the AlSi12 + TiH₂ foam pores

Slika 4: Globinski profili AES notranje površine v pori pene AlSi12 + TiH₂

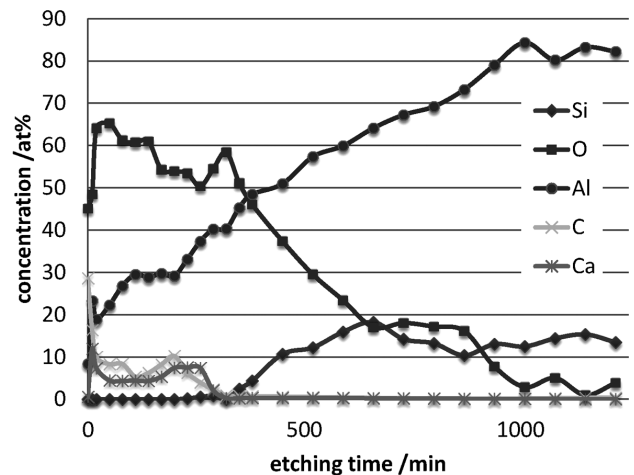


Figure 5: AES depth profiles of the interior surface of the AlSi12 + CaCO₃ foam pores

Slika 5: Globinski profili AES notranje površine v pori pene AlSi12 + CaCO₃

the oxide layer thicknesses was explained by the fact that H₂ represented a reduction atmosphere in the material, and oxygen formed on the surface of the matrix material only a thin oxide layer. Oxide was present in the material as a thin oxide layer on the surface of the powder particles. On the other hand, the CaCO₃ powder created a CO₂ atmosphere during the foaming process, so that the oxide layer on the pore surface was easily formed. Thick oxide layers enabled better stability of the foams produced by foaming agents based on CO₂. However, in both cases there was no great difference in the thicknesses of the oxide layers on the external surface.

The question appears, why CO₂-based foaming agents are not yet more widely used? And the answer is that agents releasing CO₂ need higher temperatures for the liberation of the gas and the needed temperature is approximately 100 °C above the melting point of pure aluminium¹⁰. In such a case, the matrix material is already molten and the viscosity of such material is not high enough to prevent gas escaping out of the material.

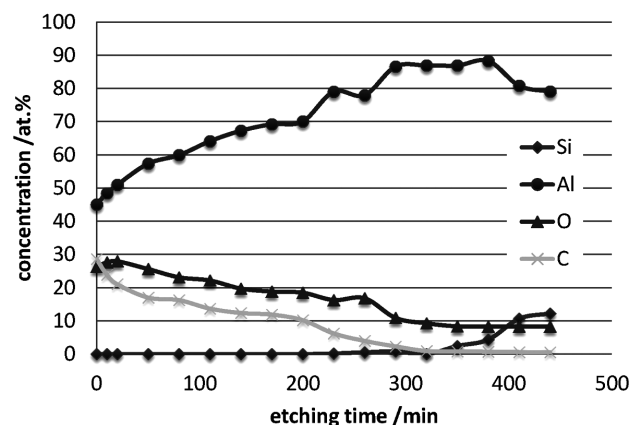


Figure 6: AES depth profiles of the exterior AlSi12 + TiH₂ foam surface

Slika 6: Globinski profili AES zunanje površine pene AlSi12 + TiH₂

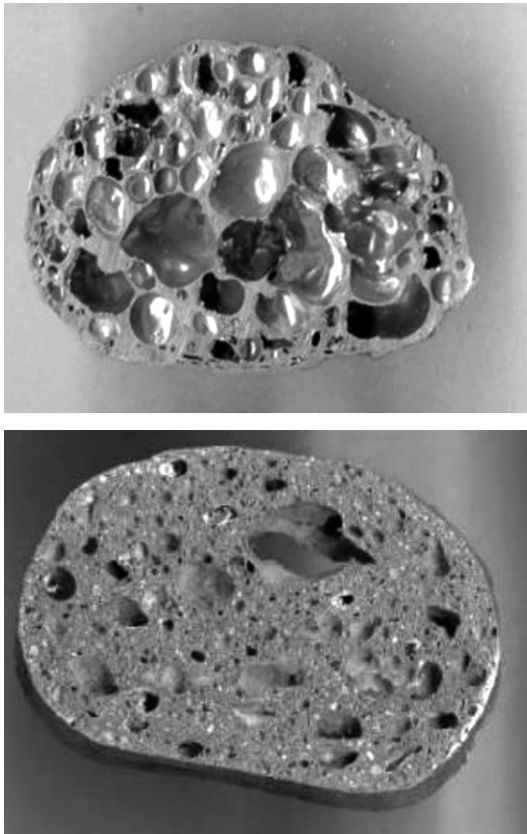


Figure 7: Images of Al foam, cut from loaf: a) AlSi12 + TiH₂ and b) AlSi12 + CaCO₃

Slika 7: Posnetka Al pene, odrezan Al hlebček: a) AlSi12 + TiH₂ in b) AlSi12 + CaCO₃

The solution is to pre-heat the precursors for some time before the foaming process begins. When foaming agents in those precursors are then exposed to an elevated temperature, decomposition has already started and the agent did not need such a long time to decompose completely, while the matrix aluminium has insufficient time to melt completely.

The sizes and shapes of the pores in the foams prepared by exposure to H₂ or CO₂ gas are different. H₂ gas forms spherical pores, while CO₂ forms smaller pores of not completely spherical shape; they are slightly elongated in the horizontal direction (**Figure 7**). However, the stability of foams is better when a CO₂-based foaming agent is used, but the pore shapes and consequently the mechanical properties in the vertical direction are better when using H₂-based foaming agents. The final application of the foamed material determines the selection of the foaming agent as well as the selection of the matrix material. On the other hand, the use of a combination of both foaming agents in a well-defined proportion can be an alternative.

4 CONCLUSIONS

CaCO₃ is a good substitute for the expensive and more widely used TiH₂ foaming agent, though some properties of foams, like the stabilization of pores, are in this case even better. Natural oxidation of the pore surfaces inside the material, as well as the oxidation of external surfaces, helps to stabilize foams so that they do not collapse after the completed foaming.

The oxide layers detected by the AES analysis on the surface of aluminium powders helped to prevent premature melting of the matrix material and enabled foaming agents to decompose before the matrix material melted and lost sufficient viscosity. Similar case is observed when using TiH₂ foaming agent that decomposes at a relatively low temperature, but the additional oxide layer on the surface of the powder delays the beginning of the decomposition as it was pre-studied and confirmed in the literature.

According to^{12,13} and from our research the conclusion can be made that a thicker oxide layer and the residuals of the ceramic solid particles from the CaCO₃ foaming agent give an opportunity to foaming agents based on CO₂ to be used more often in the production of aluminium foams.

Nevertheless, the stability of the foams is better when a CO₂-based foaming agent is used, but the pore shapes are more spherical and consequently the mechanical properties, as known from the literature, are higher when H₂-based foaming agents are used, so the use of a combination of both foaming agents in a well-defined proportion can be an alternative.

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